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11089P1 GB/ED

2. Patent application number

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3. Full name, address and postcode of the or of each applicant (underline all surnames)

Reckitt Benckiser (UK) Limited  
103 - 105 Bath Road  
Slough  
Berkshire  
SL1 3UH  
United Kingdom

Patents ADP number (if you know it)

07972136002 ✓

If the applicant is a corporate body, give the country/state of its incorporation

England

4. Title of the invention

Improvements in or relating to cleaning

5. Name of your agent (if you have one)

Elizabeth Anne Dickson

"Address for service" in the United Kingdom to which all correspondence should be sent (including the postcode)

Reckitt Benckiser plc  
Group Patents Department  
Dansom Lane  
HULL  
HU8 7DS  
UNITED KINGDOM

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07799521001 ✓

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Country

Priority application number  
(if you know it)

Date of filing

7. If this application is divided or otherwise derived from an earlier UK application, give the number and the filing date of the earlier application

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8. Is a statement of inventorship and of right to grant of a patent required in support of this request? (Answer 'Yes' if:

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I/We request the grant of a patent on the basis of this application.

11.

Signature

Date

Elizabeth Anne Dickson

7 August 2002

12. Name and daytime telephone number of Person to contact in the United Kingdom

Elizabeth Anne Dickson (01482) 582909

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## Improvements in or relating to cleaning

The present invention relates to a cleaning composition, a cleaning pack and a method of cleaning. The composition, pack and method are of particular interest in relation to the cleaning of hard surfaces, especially sanitaryware, notably lavatories.

Lavatory cleaners are purchased by customers with a view to killing germs and also removing limescale and other stains. Typically, conventional lavatory cleaners are based on sodium hypochlorite or hydrochloric acid. They may be supplied as thin liquid compositions, or as thickened compositions. Thin liquid compositions have the disadvantage of short residence time on an upright or inclined sanitaryware surface, above a water line. Thickened compositions may have a longer residence time on an upright or inclined sanitaryware surface, but may be more difficult to dispense from a typical squeeze pack.

Compositions have been disclosed in which there is a first liquid containing sodium chlorite, stabilised under alkaline conditions, and a second, acidic, liquid, the acid being present in an amount in excess of that needed to neutralise the first liquid. The two liquids are mixed at the point of use to form a cleaning composition. The sodium chlorite is then in an acidic environment, and releases chlorine dioxide. An example of such a composition is described in WO 98/57544. However, in such examples the problem of using thin, non-clinging liquids or viscous, difficult to impel liquids still exists.

In accordance with a first aspect of the invention there is provided a hard surface cleaner comprising a first precursor composition, being a thin alkaline liquid which contains a compound which generates chlorine dioxide under acidic conditions but not under alkaline conditions, and a  
5 second precursor composition, being a thin acidic liquid, wherein on mixing the first and second precursor compositions the resulting admixed cleaning composition is acidic, causing the compound to generate chlorine dioxide,  
10 and is more viscous than the first precursor composition and more viscous than the second precursor composition.

Although chlorine dioxide is generated on admixture, preferably the cleaning composition is formulated to be  
15 non-foaming.

Preferably the first precursor composition is an aqueous composition.

20 Preferably the second precursor composition is an aqueous composition.

Preferably the cleaning composition does not generate chlorine gas.

25

Preferably the cleaning composition comprises a colour change agent. Initially, on mixing of the two precursor compounds, the resulting admixed cleaning composition is of a first appearance but changes to a second appearance,  
30 of different colour, preferably over a short period of time, preferably less than 5 minutes, and more preferably less than 2 minutes. The colour change may be promoted by any chemical change at work. For example it may be caused

by the breakdown of the compound which generates chlorine dioxide. It may be promoted by the evolution of chlorine dioxide. It may be promoted by the pH change.

- 5 A colour change agent may be contained within the first precursor composition, or the second precursor composition, or both.

10 Preferably, a colour change agent is contained within the second precursor composition and employs a colour change agent which is oxidised by the chlorine dioxide evolved. Preferably it comprises a dye which is oxidised and a dye which is resistant to oxidation, so that the change is not from coloured to colourless, but from a first colour to a  
15 second colour. An example of a suitable oxidisable dye is acid blue. An example of a dye resistant to oxidation is methylene blue.

By "thin" we mean that the viscosity of such a liquid is  
20 like that of water, or not very much higher than that. Suitably a thin liquid herein has a viscosity of less than 100 cps, preferably less than 50 cps, and more preferably less than 20 cps. Especially favoured thin liquids have a viscosity less than 10 cps.

25

Suitably a thin liquid herein has a viscosity of at least 0.5 cps, preferably at least 1 cps, most preferably at least 2 cps.

30 Preferably the viscosity of the admixed cleaning composition is at least 150 cps, more preferably at least 400 cps, and most preferably at least 500 cps. It is

especially preferred that the admixed cleaning composition is of a viscosity greater than 700 cps.

Preferably the viscosity of the admixed cleaning composition does not exceed 4,000 cps. Typically, it does not exceed 2,000 cps, or, preferably, 1000 cps.

For reference purposes, when we give values herein for viscosity they are measured at 20°C using a Brookfield viscometer, Spindle LV 2, 60 rpm.

Preferably, then, on mixing of the two precursor compositions the resulting admixed cleaning composition thickens, so that it has a longer residence time on an upright or inclined hard surface. Preferably its viscosity and/or adherence are such that the admixed cleaning composition, or at least a proportion of it, clings to the surface until flushed therefrom.

Preferably the first precursor composition and/or the second precursor composition contains a surfactant which thickens the admixed composition, when admixture takes place. In order not to thicken the precursor composition in which it initially resides, the surfactant should be responsive to the other precursor composition. Most conveniently the surfactant is responsive to pH. Since the second precursor composition and the admixed cleaning composition are both acidic, the surfactant is preferably comprised within the first precursor composition and provides the thickening effect when it is exposed to the acid in the second precursor composition.

Thus, preferably the surfactant selected is one which provides an increase in viscosity when it undergoes a change from an alkaline to an acidic environment.

5 Preferred surfactants for use in the present invention are anionic surfactants which provide any of the effects stated above. Especially preferred are alkyl sulphonate, or more preferably, alkyl sulphate surfactants, especially C<sub>8-20</sub> (preferably linear) alkyl sulphate surfactants.

10

Preferred alkyl sulphate surfactants are alkali metal salts, especially sodium salts.

Preferred anionic surfactants are alkoxyated, preferably  
15 by 1-8, more preferably 1-4, especially 1 or 2, alkoxy groups per anionic surfactant molecule, on average. Preferred alkoxy groups are ethoxy groups, although other alkoxy groups, notably propoxy groups, are not excluded.

20 Preferably the first precursor liquid contains, as the compound which generates chlorine dioxide under acidic conditions, an alkali metal chlorate or, preferably, an alkali metal chlorite. Sodium salts are preferred.

25 Optionally the first and/or second precursor liquid contains, in addition to materials already mentioned, a water-soluble metal salt. The cation is preferably an alkali metal ion, most preferably a sodium ion. The anion is preferably derived from a mineral acid. Most  
30 preferably it is a halide ion, especially a chloride ion.

Preferably the pH of the first precursor composition is at least 8, preferably at least 10, and more preferably at



least 12. The compound which generates chlorine dioxide under acidic conditions, contained with the first precursor composition, may itself be a base but the first composition preferably contains an additional base, suitably an alkali metal base, for example a hydroxide. Sodium hydroxide is favoured for this purpose.

When an additional base is provided in the first liquid precursor, it is suitably provided in an amount of at least 0.05% w/w, preferably at least 0.1% w/w, based on the total cleaning composition.

When an additional base is provided in the first liquid precursor, it is suitably provided in an amount up to 1% w/w, preferably up to 0.5% w/w, based on the total cleaning composition.

Preferably the second precursor liquid has a pH of up to 5, more preferably up to 3.5, and most preferably up to 2.5. Suitably the acidifying compound is a mineral acid, preferably hydrochloric acid.

The pH of the admixed cleaning composition will exceed the pH of the second precursor composition, due to the alkalinity of the first precursor composition. The pH of the admixed cleaning composition is preferably no more than 0.5 pH units higher than that of the second precursor or liquid, and so is preferably of pH up to 5.5, more preferably up to 4, and most preferably up to 3.

30

Unless otherwise stated pH values are stated herein with reference to a composition diluted at a ratio of 1 part of composition to 100 parts of deionised water, by volume.

The composition may contain other ingredients including additional surfactants, fragrances, stabilisers and buffers.

5

Additional surfactants include nonionic and amphoteric (or zwitterionic) surfactants. When present the amount of such additional surfactant, or of such additional surfactants in total, is preferably up to 10% w/w, more  
10 preferably up 5% w/w, of the total cleaning composition.

One class of nonionic surfactants which may be used in the present invention is alkoxyated alcohols, particularly alkoxyated fatty alcohols. These include ethoxylated and  
15 propoxylated fatty alcohols, as well as ethoxylated and propoxylated alkyl phenols, both having alkyl groups of from 7 to 16, more preferably 8 to 13 carbon chains in length. Examples include linear alcohol ethoxylates and secondary alcohol ethoxylates.

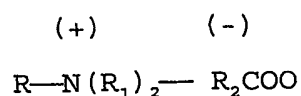
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Another class of non-ionic surfactants that may be used is sorbitan esters of fatty acids, typically of fatty acids having from 10 to 24 carbon atoms, for example sorbitan mono oleate.

25

Amphoteric surfactants which may be used in the present invention including amphoteric betaine surfactant compounds having the following general formula:

30



wherein R is a hydrophobic group which is an alkyl group containing from 10 to 22 carbon atoms, preferably from 12 to 18 carbon atoms, an alkylaryl or arylalkyl group containing a similar number of carbon atoms with a benzene ring being treated as equivalent to about 2 carbon atoms, and similar structures interrupted by amido or either linkages; each  $R_1$  is an alkyl group containing from 1 to 3 carbon atoms; and  $R_2$  is an alkylene group containing from 1 to 6 carbon atoms.

10

A preferred amphoteric surfactant includes an alkyl amino betaine or an alkyl amido betaine.

Suitable amphoteric surfactants also include cocoamides, preferably polyoxyethylene-3-cocoamide.

15

Suitable amphoteric surfactants also include imidazoline surfactants, for example sodium capryloamphopropionate.

Suitable amphoteric surfactants include lactamide surfactants, for example Lactamide MEA.

20

Preferably the compound which generates chlorine dioxide under acidic conditions is present in an amount up to 2% w/w of the total cleaning composition, preferably up to 1% w/w of the total cleaning composition. Preferably it is present in an amount of at least 0.02% w/w of the total cleaning composition, preferably at least 0.1% w/w of the total cleaning composition.

25

Preferably a surfactant responsible for the thickening effect, when present, is present in an amount of at least 0.5% w/w of the total cleaning composition, preferably at

30

least 1% w/w of the total cleaning composition, and most preferably at least 1.5% w/w of the total cleaning composition. Preferably such a surfactant is present in an amount up to 6% w/w of the total cleaning composition, preferably up to 4% w/w, of the total cleaning composition.

Preferably the second precursor liquid contains an acid in an amount up to 15% w/w of the total cleaning composition, preferably up to 10% w/w of the total cleaning composition, and most preferably up to 7% w/w of the total cleaning composition. Preferably such an acid is present in an amount at least 2% w/w of the total cleaning composition, more preferably at least 4% w/w of the total cleaning composition; but in all cases in an amount in excess of that required to neutralise an equal volume of the first precursor composition.

Preferably the normality of an acid present in the second liquid composition is at least 1M, more preferably at least 1.5M, preferably at least 2M. Preferably the normality of the acid in the admixed cleaning composition is at least 0.5M, more preferably at least 0.8M.

Preferably water from the first liquid precursor provides at least 20% w/w of the total cleaning composition, more preferably at least 24% w/w, most preferably at least 28% w/w.

Preferably water from the first liquid precursor provides up to 46% w/w of, more preferably up to 42% w/w, most preferably up to 38% w/w of the total cleaning composition.

Preferably water from the second liquid precursor provides at least 20% w/w of the total cleaning composition, more preferably at least 24% w/w, most preferably at least 28% w/w.

Preferably water from the second liquid precursor provides up to 45% w/w of, more preferably up to 40% w/w, most preferably up to 35%.

10

Preferably water comprises at least 40% w/w of the admixed cleaning composition, more preferably at least 48%, most preferably at least 56% w/w. Especially preferred are admixed compositions at least 60% water.

15

Preferably water comprises up to 90% w/w of the admixed cleaning composition, more preferably up to 82% w/w, most preferably up to 74% w/w. Especially preferred are admixed compositions comprising up to 70% water.

20

Preferably there is provided a hard surface cleaner as defined herein, which comprises (% w/w values stated with reference to the total cleaning composition):

25 0.02-2% w/w, preferably 0.1-1% w/w, of a compound which under acid conditions generates chlorine dioxide, provided substantially wholly via the first precursor liquid;

0.5-5% w/w, preferably 1.5-4% w/w, of an alkali metal, alkoxyated C<sub>8-20</sub> alkyl sulphate surfactant which thickens on being acidified, provided partially or substantially wholly via the first precursor liquid;

30

20-46% w/w, preferably 28-40% w/w, water, provided via the first precursor liquid;

5 an alkali metal base, provided substantially wholly via the first precursor liquid, in an amount such as to make the pH thereof at least 12;

20-46% w/w, preferably 28-40% w/w, water, provided via the second precursor liquid;

10

an acid, provided substantially wholly via the second precursor liquid in an amount such as to make the pH thereof not more than 5 and the pH of the admixed cleaning composition not more than 0.5 higher than the pH of the  
15 second precursor liquid;

wherein the viscosity (as measured herein) of the first precursor liquid is in the range 0.1-100 cps, preferably 1-10 cps, the viscosity of the second precursor liquid is  
20 in the range 0.1-100 cps, preferably 1-10 cps, and the viscosity of the admixed composition is in the range 150-1400 cps, preferably 500-1000 cps.

In accordance a second aspect of the invention there is  
25 provided a hard surface cleaning pack comprising a first chamber containing the first precursor composition, and a second chamber containing the second precursor composition, wherein the chambers are adapted for simultaneous, separate dispensing of the first and second  
30 precursor compositions with downstream mixing thereof, the first and second precursor compounds being in accordance with the first aspect of the invention.

The mixing of the first and second precursor compositions can be within a mixing zone adjacent to the outlets of the first and second chambers. However with this kind of arrangement steps may need to be taken to avoid contamination of the contents of the chambers. Therefore  
5 in a preferred embodiment the pack is designed to issue separate first and second precursor compositions which mix substantially only when contacting the hard surface.

10 Containers/dispensers suitable for use in this aspect of the invention are well known. Examples include twin-pack trigger sprays having a mechanical pumping action and side-by-side twin squeeze bottle chambers having simple narrowed nozzle outlets.

15 In accordance with a third aspect of the invention there is provided a method of cleaning a hard surface, the method comprising use of a cleaning composition which is an in situ mixture of the first precursor composition and  
20 the second precursor composition, as defined above in relation to the first aspect of the invention.

Preferably the hard surface is an inclined or upright surface.

25 Preferably the hard surface is a sanitaryware surface.

Preferably the hard surface is a hard surface of a lavatory, including of a lavatory bowl or of a urinal.

30 Preferably the method employs 30-70 parts by weight of the first precursor composition, more preferably 40-60 and most preferably 45-55, with the second precursor

10 It will be appreciated that the present invention offers many benefits to the consumer. Thin liquids are employed, and so are easier to dispense than a pre-thickened cleaning composition. The thin liquids become thick on admixture and so attain prolonged residence time, on an upright or inclined surface. On admixture chlorine dioxide is generated. Chlorine dioxide is an effective bleaching and sanitising agent, of sufficiently low odour that it can be masked by a fragrance. The admixed composition is acidic and is able to combat limescale.

20 The colour change agent, when present, gives the consumer a visual indication of these actions. Since in preferred embodiments the colour change agent is responsive directly to the presence of chlorine dioxide, whose evolution is triggered by acid conditions, the colour change can be no mere marketing artefact, but a direct indication of the chemistry which is at work.

25

The invention will now be further described, by way of illustration, with reference to the following examples. .

30 In these examples the following materials are used.

EMPICOL ESB - Sodium C(10-19)-(ethoxy)<sub>2</sub>-sulphate,  
from Huntsman



- 
- EMPICOL ESA - Sodium C(10-19)-ethoxy-sulphate, from  
Huntsman
- 5 Acid Blue 182 - oxidisable blue dye, from Clariant
- Methylene blue - oxidation-resistant blue dye, from  
Orion Dyes, NJ
- 10 EMPICOL ESA and EMPICOL ESB are trade marks.

Example 1 - formulation

15 This example is of a dual liquid system to be jetted under  
the toilet rim via a two compartment bottle. Equal  
volumes of each liquid are mixed on delivery, at the hard  
surface.

20 **Liquid A** is a clear thin liquid containing an unactivated  
oxidising agent, an alkaline stabilising agent, anionic  
surfactant, and fragrance, in water. pH at 20°C is 12.55

**Liquid B** is a thin blue liquid containing mineral acid and  
blue dyes in hard water. pH at 20°C is 1.83.

25

The liquids are prepared by simple blending.

Full composition information is given below.

Ingredient	LIQUID A Function	% w/w Liquid A	% w/w Total
Deionised water	Dilutant	59.50	29.75
Sodium Hydroxide solution 4 % w/w, in water	Stabilising Agent	10.00	5.00
Sodium Chlorite 80% w/w, in water	Oxidising Agent	0.50	0.25
Sodium Alkyl ethoxy sulphate 2EO (25.5% soln, in water)	Anionic Surfactant/Thickener	20.00	10.00
Fragrance pre-mix	Fragrance	10.00	
		100.00	
<u>Fragrance pre-mix</u>			
Commercial fragrance	Fragrance	2.00	0.100
Sodium Alkyl ethoxy sulphate 2EO 25.5 % soln in water)	Anionic Surfactant	50.00	2.500
Deionised Water	Dilutant	48.00	2.400
		100.00	50.000

Ingredient	LIQUID B Function	% w/w Liquid B	% w/w Total
Hard water	Dilutant	65.00	32.50
Hydrochloric Acid 37% w/w, in water	Acid	25.00	12.50
Dye pre-mix	Dye	10.00	
		100.00	
<u>Dye pre-mix</u>			
Acid Blue 182	Readily Oxidised Dye	0.4000	0.020
Methylene Blue	Oxidation Resistant Dye	0.0250	0.001
Hard Water	Dilutant	99.5750	4.979
		100.0000	50.000

The two thin liquids mix to form a dark blue thick liquid, whose pH at 20°C is 1.96. The acid activates the oxidising agent liberating chlorine dioxide. The acid

removes limescale. The surfactant causes the substantial thickening, and aids surface activity.

5 A colour change from blue to indigo then to green occurs over approximately 1 minute. On standing overnight the colour returns to pale blue.

pH values stated in this specification are of a 1/100<sup>th</sup> dilution of a sample in deionised water.

10

#### Example 1 - viscosity

The viscosity of Liquid A and Liquid B and of the admixed cleaning composition were tested at 20°C using a  
15 Brookfield viscometer, Spindle LV 2, at 60 rpm.

The viscosity of Liquid A was 5 cps.

The viscosity of Liquid B was 4 cps.

20

The viscosity of the admixed cleaning composition was 774 cps.

25 The practical effect of this viscosity increase is that whereas either Liquid A or Liquid B alone quickly run off an inclined surface of a lavatory bowl, when they are mixed the resulting thicker liquid "curtains" evenly and slowly down the surface. A proportion runs into the water in the lavatory bowl, where some of it "pools" at the  
30 bottom of the water, in the region where limescale may typically be most prevalent. The remainder clings to the surface, above the water line. After one flush of the

lavatory there is significant foaming. Residual foam is still apparent after five flushes.

Example 1 - limescale removal

5

A marble cube (approximately 25g in weight) was weighed then added to water (2 litres) in a lavatory bowl. 25 g of Liquid A and 25 g of Liquid B were mixed, as they were applied to the region under the rim of a lavatory bowl.

10 The cube was left in contact with the liquid for 6 hours. At the end of this time period the cube was removed with plastic forceps. It was rinsed with tap water, and then deionised water, before drying off any surface water with a paper towel. The marble was further dried in a  
15 microwave, allowed to cool, and reweighed.

In a test of limescale combating power of the neat cleaning composition 50 g of the neat cleaning composition was placed in a 100 ml beaker. A pre-weighed marble cube  
20 was added. It was completely immersed in the cleaning composition. In this test the test periods were 20 minutes, 3 hours and 6 hours. Drying/weighing were as described above.

25 The results for Example 1 are set out below. As an aid to the interpretation of the figures for Example 1, the following commercial products were tested in the same way.

DOMESTOS OX

30 HARPIC 100% Limescale Remover (HARPIC LSR)  
HARPIC Clear Liquid Bleach (HARPIC CSR)  
DOMESTOS Bleach

Diluted:

	Initial (g)	After 6h (g)	Weight lost (g) in 6h
Example 1	25.0673	23.5559	1.5114
DOMESTOS OX	24.9239	24.1159	0.808
HARPIC LSR	24.7797	23.8542	0.9255
HARPIC CSR	24.923	24.9277	None
DOMESTOS bleach	24.6959	24.6945	0.0014

Neat:

5

	Initial (g)	20 mins (g)	3 hours (g)	6 hours (g)	Weight lost		
					20 mins (g)	3 hours (g)	6 hours (g)
Example 1	22.8308	22.6879	21.8477	21.23	0.1429	0.9831	1.6008
DOMESTOS OX	19.0143	18.9864	18.6069	18.3977	0.0279	0.4074	0.6166
HARPIC LSR	24.5556	24.1326	22.4635	22.0644	0.423	2.0921	2.4912
HARPIC CSR	20.2957	20.2983	20.2979	20.2937	-	-	0.002
DOMESTOS bleach	19.0184	19.0181	19.0152	19.0145	0.0003	0.0032	0.0039

It will be seen that the limescale removal performance of Example 1 is good to very good.

Example 1 - bleaching

The Example 1 cleaning composition, admixed just prior to  
5 testing, was tested for its bleaching power compared with  
HARPIC CSR, HARPIC LSR, DOMESTOS bleach and DOMESTOS OX.

## 1. Preparation of test solution No. 1:

10 sodium metasilicate (2.0 g) was added to deionised  
water (46.5 ml). After dissolution of the solid,  
sodium bicarbonate (1.5 g) was added and stirring was  
continued until all of the solid had dissolved.

## 2. Preparation of test solution No. 2:

15 Calcium chloride (1.0 g) was dissolved in deionised  
water (36.0 ml). Magnesium chloride hexahydrate (0.50  
g) was then added to the solution. A food grade blue  
dye (75 mg) and ethanol (12.5 ml) were added and the  
solution stirred until a uniform solution was  
20 obtained.

3. A clean white ceramic tile was placed on a balance and  
tared. Solution No. 2 (1.8g) was added dropwise to  
the tile. Solution No. 1 (2.6g) was then added to the  
25 tile. The solutions were mixed and evenly distributed  
about the surface of the tile using a toothbrush. The  
prepared tile was left to dry overnight at ambient  
temperature.

30 Testing of bleaching above the water line was as follows.

A few drops of each formulation to be tested (including  
deionised water for comparison) were placed in a spaced-

- apart - array - on - the - surface - of - the - prepared - tile. . . . The . . . tiles were left for 1 hour. The formulations were removed from the tiles using clean tissues (tissues were placed on surface of tile to soak up formulation without washing the dye off the stain, and were carefully removed).

Visual assessment was then made. The Example 1 cleaning composition showed a similar level of bleaching activity to HARPIC CSR and DOMESTOS. It showed better bleaching activity than HARPIC LSR and DOMESTOS OX. The latter two products did not show detectable bleaching in this test. They did show some apparent dissolution of the dye and left prominent "tide marks".

#### 15 Example 1 - biocidal efficacy

The Example 1 cleaning composition was tested for biocidal efficacy. For reference purposes HARPIC CSR and HARPIC LSR were subjected to the same tests.

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The samples were tested using the BS EN1276 suspension test for bactericidal activity, was adapted to provide a more quantitative assessment of efficacy. Samples were tested both neat and diluted to simulate action above and below the water line respectively, in a 2-litre toilet bowl. Activity above the water line was tested after a 2 minute contact time to coincide with the colour change associated with the production of chlorine dioxide. Activity below the water line was tested after a 5 minute contact time. The test was performed against *Staphylococcus aureus* and *Pseudomonas aeruginosa* in the presence of organic soil (0.3% bovine albumin in-test).

### Neat Samples

Neat samples were evaluated for antimicrobial activity as follows:

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- 1ml bacterial suspension ( $10^8$ cfu/ml) was transferred to a flask containing 1ml of a 3% BA suspension (dirty conditions).
- The culture/soil mix was vortexed and then shaken in a water bath (20°C) for 2 minutes.
- To the culture/soil mix 4ml of each of Liquid A and Liquid B of the Example 1 cleaning composition were added simultaneously using a syringe. For the single component liquids (benchmarks), 8ml was added.
- 2 minutes after the addition of the sample to the culture/soil mix, a 100µl aliquot of the test mixture was removed and transferred to 9.9ml neutralising medium.
- After a neutralisation period of 5 minutes, the sample was serially diluted and used to prepare pour plates which were subsequently incubated at  $36^\circ\text{C} \pm 1^\circ\text{C}$  for 48 hours before enumerating surviving bacteria.
- The test was performed in duplicate.

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### 25 Diluted Samples

Diluted samples were prepared in 400ml of sterile hard water (300ppm  $\text{CaCO}_3$  in-test) as follows:

- For the Example 1 cleaning composition, 4.8ml each of liquids A and B were mixed in a sterile pot immediately prior to testing to ensure adequate thickening before addition to the test beaker. Each sample was then added

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to the beaker by pouring around the rim; to simulate application of the product to the toilet bowl.

The composition was left to pool in the hard water for 5 minutes following application to the beaker rim.

After the 5 minute pooling time, the samples were mixed gently using a magnetic stirrer and immediately evaluated for antimicrobial activity exactly as described above for the neat samples but incorporating the following changes:

- 8ml of the diluted product was added to the soil/culture mixture
- A contact time of 5 minutes was used rather than 2 minutes
- Neutralisation was performed by removing a 1ml aliquot of the test mixture and transferring it to 9.0ml neutralising medium
- The test was performed in triplicate

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Neutralisation validation was performed to ensure that all samples were adequately neutralised during the test.

### Results

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The Microbiocidal Effect (ME) due to the action of the test sample after 2 minutes (neat) or 5 minutes (diluted), was then calculated using the following equation:-

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$$ME = \log N_c - \log N_d$$

$N_c$  = number of cfu/ml added at time 0

$N_d$  = number of cfu/ml for the test mixture after the action of the test sample

The results are set out in the table below

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Test Sample	ME Value (log cfu/ml)					
	<i>Staphylococcus aureus</i>			<i>Pseudomonas aeruginosa</i>		
Example 1 Neat	>5.2	>5.2		>4.7	>4.7	
Example 1 cleaning composition Diluted	>5.1	>5.1	>5.2	>4.7	>4.7	>4.9

The results show that exposure of bacteria to the Example 1 cleaning composition in a standard BS EN1276 suspension test produced considerable reductions in viability.

Reductions in excess of 5.0 log cfu/ml were achieved against *Staphylococcus aureus* for the neat and diluted Example 1 cleaning composition, in the presence of organic soil and hard water after a 2 minute (neat) or 5 minute (diluted) contact time.

Reductions in excess of 4.5 log cfu/ml were achieved against *Pseudomonas aeruginosa* for the neat and diluted Example 1 cleaning composition in the presence of organic soil and hard water after a 2 minute (neat) or 5 minute (diluted) contact time.

In terms of antimicrobial activity, in comparison with commercial products, the Example 1 cleaning composition gave comparable activity in the tests of action when neat, relevant to action above the water line, and was superior

----- in the tests when diluted, relevant to action below the ---  
water line.

Example 2

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This is of a dual liquid system to be added to the toilet rim via a two compartment bottle. Equal volumes of each liquid are mixed on delivery.

- 10 Liquid A consisted of an unactivated oxidising agent, stabilising agent, anionic surfactant and fragrance.

Liquid B consisted of a dilute mineral acid and blue dyes.

- 15 Full compositional information is given below.
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Ingredient	LIQUID A Function	% w/w Liquid A	% w/w Total
Deionised water	Dilutant	62.50	31.25
Sodium Hydroxide solution 4% w/w, in water	Stabilising Agent	10.00	5.00
Sodium Chlorite 80% w/w, in water	Oxidising Agent	0.50	0.25
Sodium Alkyl ethoxy sulphate 1EO (25% soln, in water)	Anionic Surfactant/Thickener	17.00	8.50
Fragrance pre-mix	Fragrance	10.00	
		<u>100.00</u>	
<u>Fragrance pre-mix</u>			
Commercial fragrance	Fragrance	2.00	0.100
Sodium Alkyl ethoxy sulphate 1EO (25% soln, in water)	Anionic Surfactant	50.00	2.500
Deionised Water	Dilutant	48.00	2.400
		<u>100.00</u>	<u>50.000</u>

Ingredient	LIQUID B Function	% w/w Liquid B	% w/w Total
Hard water	Dilutant	65.00	32.50
Hydrochloric Acid 37% w/w, in water	Acid	25.00	12.50
Dye pre-mix	Dye	10.00	
		<u>100.00</u>	
<u>Dye pre-mix</u>			
Acid Blue 182	Readily Oxidised Dye	0.4000	0.020
Methylene Blue	Oxidation Resistant Dye	0.0250	0.001
Hard Water	Dilutant	99.5750	4.979
		<u>100.0000</u>	<u>50.000</u>

The two thin liquids mix to form a dark blue thick liquid. The acid activates the oxidising agent liberating chlorine dioxide. The acid removes limescale. The surfactant causes the substantial thickening, and aids surface  
5 action.

A colour change from blue to indigo then to green occurs over approximately 1 minute. On standing overnight the colour becomes pale blue.

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### Example 3

This is of a dual liquid system to be added to the toilet rim via a two compartment bottle. Equal volumes of each  
15 liquid are mixed on delivery.

Liquid A consisted of an unactivated oxidising agent, stabilising agent and anionic surfactant.

20 Liquid B consisted of a dilute mineral acid, anionic surfactant, blue dyes and fragrance.

Full compositional information is given below.

Ingredient	LIQUID A Function	% w/w Liquid A	% w/w Total
Deionised water	Dilutant	71.29	35.65
Sodium Hydroxide solution 4% w/w, in water	Stabilising Agent	10.00	5.00
Sodium Chlorite 80% w/w, in water	Oxidising Agent	0.50	0.25
Sodium Alkyl ethoxy sulphate 2EO (28% soln, in water)	Anionic Surfactant/Thickener	18.21	9.11
		<u>100.00</u>	

Ingredient	LIQUID B Function	% w/w Liquid B	% w/w Total
Hard water	Dilutant	65.00	32.50
Hydrochloric Acid 37% w/w, in water	Acid	25.00	12.50
Dye and fragrance pre-mix	Dye and fragrance	10.00	
		<u>100.00</u>	

Dye and fragrance pre-mix

Acid Blue 182	Readily Oxidised Dye	0.4000	0.020
Methylene Blue	Oxidation Resistant Dye	0.0250	0.001
Commercial fragrance	Fragrance	2.00	0.100
Sodium alkyl ethoxy sulphate 2EO (28% soln, in water)	Anionic surfactant	45.5400	2.277
Hard Water	Dilutant	<u>52.0350</u>	<u>2.602</u>
		<u>100.0000</u>	<u>50.000</u>

The two thin liquids mix to form a dark blue thick liquid. The acid activates the oxidising agent liberating chlorine dioxide. The acid removes limescale. The surfactant causes the substantial thickening, and aids surface activity.

A colour change from blue to indigo then to green occurs over approximately 1 minute. On standing overnight the colour returned to pale blue.

CLAIMS

1. A hard surface cleaner comprising a first precursor composition, being a thin alkaline liquid which contains a compound which generates chlorine dioxide under acidic conditions but not under alkaline conditions, and a second precursor composition, being a thin acidic liquid, wherein on mixing the first and second precursor compositions the resulting admixed cleaning composition is acidic, causing the compound to generate chlorine dioxide, and is more viscous than the first precursor composition and more viscous than the second precursor composition.
2. A cleaning composition cleaner as claimed in claim 1, wherein the viscosity of the first precursor composition is in the range 0.5-100 cps, the viscosity of the second precursor composition is in the range 0.5-100 cps and the viscosity of the cleaning composition produced by admixture thereof is 150-4000 cps.
3. A cleaning composition as claimed in claim 1 or 2 wherein the first precursor composition is of pH at least 8, and comprises an alkali metal chlorite stabilised in aqueous solution by an additional base.
4. A cleaning composition as claimed in any preceding claim, wherein the second precursor composition is of pH not more than 5, and comprises a mineral acid in aqueous solution.
5. A cleaning composition as claimed in any preceding claim, wherein the pH of the cleaning composition, formed



0.02-2% w/w, preferably 0.1-1% w/w, of a compound which under acid conditions generates chlorine dioxide, provided substantially wholly via the first precursor liquid;

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0.5-5% w/w, preferably 1.5-4% w/w, of an alkali metal, alkoxyated C<sub>8-20</sub> alkyl sulphate surfactant which thickens on being acidified, provided partially or substantially wholly via the first precursor liquid;

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20-46% w/w, preferably 28-40% w/w, water, provided via the first precursor liquid;

an alkali metal base, provided substantially wholly via the first precursor liquid, in an amount such as to make the pH thereof at least 12;

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20-46% w/w, preferably 28-40% w/w, water, provided via the second precursor liquid;

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an acid, provided substantially wholly via the second precursor liquid, in an amount such as to make the pH thereof not more than 5 and the pH of the admixed cleaning composition not more than 0.5 higher than the pH of the second precursor liquid;

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wherein the viscosity (as measured herein) of the first precursor liquid is in the range 0.1-100 cps, preferably 1-10 cps, the viscosity of the second precursor liquid is in the range 0.1-100 cps, preferably 1-10 cps, and the viscosity of the admixed composition is in the range 150-1400 cps, preferably 500-1000 cps.

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13. A hard surface cleaning pack comprising a first chamber containing a first precursor composition, being a thin alkaline liquid which contains a compound which generates chlorine dioxide under acidic conditions but not under alkaline conditions, and a second chamber containing the second precursor composition, being a thin acidic liquid, wherein the chambers are adapted for simultaneous, separate, dispensing of the first and second precursor compositions with downstream mixing thereof, wherein on mixing the first and second precursor compositions the resulting admixed cleaning composition is acidic, causing the compound to generate chlorine dioxide, and is more viscous than both first precursor compositions.

14. A method of cleaning a hard surface, the method comprising use of a cleaning composition which is an in situ mixture of the first precursor composition and the second precursor composition, as defined in any of claims 1 to 12.

15. A method as claimed in claim 14, comprising the admixture of the first and second precursor compositions, the first precursor composition constituting 30-70 parts by weight of the cleaning composition and the second precursor composition constituting the balance.

16. A cleaning composition, cleaning pack or cleaning method, each substantially as hereinbefore described with particular reference to the examples.

ABSTRACT

## Improvements in or relating to cleaning

- 5 A two-liquid hard surface cleaner contains a thin alkaline liquid and a thin acidic liquid. Stabilised in the alkaline liquid is a compound which generates chlorine dioxide only under acidic conditions. An example is sodium chlorite. The liquids are mixed to form a thick cleaning
- 10 composition. Ethoxylated alkyl sulphate surfactants have been found to be highly effective in achieving this effect of thickening on mixing. The thickened cleaning composition is acidic and causes the evolution of chlorine dioxide, which has a sterilising and bleaching effect.
- 15 The resulting acidic liquid is also able to degrade limescale.

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